**Vocabulary: Equilibrium and Pressure**



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* Dalton’s law – a law stating that the total *pressure* exerted by a mixture of gases is equal to the sum of the *partial pressures* of the individual gases.
	+ For example, if an air tank contains oxygen with a partial pressure of 2.0 MPa (megapascals) and nitrogen with a partial pressure of 3.0 MPa, the total pressure is 5.0 MPa.
* Le Châtelier’s principle – a principle stating that a chemical equilibrium will tend to adjust to counteract any imposed changes.
	+ For example, if the pressure on an equilibrium gas mixture is increased (thus reducing volume), the equilibrium will shift in favor of the side with fewer gas molecules.
* Partial pressure – the pressure exerted by a single gas in a mixture of gases.
	+ The partial pressure of a gas in a mixture is equal to the pressure the gas would exert by itself.
* Pressure – a force exerted on a given area.
	+ Gases exert pressure on the walls of their container.
	+ In the *Equilibrium and Pressure* Gizmo™, pressure is measured in megapascals (MPa). One megapascal is about ten times atmospheric pressure (1.0 atm = 0.101325 MPa).