**Vocabulary: Melting Points**

dictionary2

**Vocabulary**

* Boiling point – the temperature at which boiling occurs.
  + The boiling point is a different temperature for every chemical.
* Covalent bond – a chemical bond in which atoms share a pair of valence electrons.
  + Covalent bonds occur between two nonmetals.
* Intermolecular forces – forces that hold together different molecules. Examples of intermolecular forces are hydrogen bonds, dipole forces, and London dispersion forces.
  + Hydrogen bonds exist in compounds that have hydrogen bonded to oxygen, fluorine, or nitrogen. Water and ammonia are examples of compounds that have hydrogen bonds.
  + Dipole forces exist in polar compounds like hydrogen sulfide.
  + London dispersion forces exist between molecules of every molecular compound and are strongest in large molecules like the molecules in paraffin wax.
  + Forces between different atoms or ions are called interatomic or interionic forces. These forces exist in chemicals that are not molecular covalent compounds.
* Ionic bond – a bond formed by the attraction between two oppositely charged ions.
  + Ionic bonds form when an anion has taken one or more electrons from a cation.
  + Ionic bonds occur between metals and nonmetals.
* Melting point – the temperature at which a solid turns into a liquid.
  + The melting point is a different temperature for every chemical.
* Metallic bond – a bond formed by each atom donating one or more electrons into the “sea of electrons.”
  + The sea of electrons forms around the positively charged metal atoms.
  + Atoms leaving a metallic bond must remove one or more electrons from the electron sea, which requires a lot of energy. Metallic bonds are very strong.
* Molecular solid – a solid substance composed of individual molecules that are held together by intermolecular forces. Common examples include ice and sugar.
  + The covalent bonds holding the atoms together in each molecule are stronger than the intermolecular forces holding the molecules together.
* Network solid – a solid substance composed of a large, covalently-bonded crystal. Graphite, diamond, and silica are common examples.
* Salt – a substance or chemical formed from a metal and a nonmetal. Salts are composed of two or more oppositely charged ions held together by strong electrostatic forces (ionic bonds).
* Smoke – solid particles suspended in a gas.

* Sublimation – the transition from a solid to a gas without going through the liquid phase.
  + At normal atmospheric temperature and pressure, carbon dioxide sublimes from a solid to a gas.
* Sublimation point – the temperature at which a solid turns into a gas.
* Transition point – the temperature at which a substance changes phase or state.
  + Transition points include melting points, boiling points, and sublimation points.