**Vocabulary: Polarity and Intermolecular Forces**



**Vocabulary**

* Dipole – a molecule containing two “poles,” or regions of opposite charge.
	+ A polar molecule contains a permanent dipole.
* Dipole-dipole force – an intermolecular force of attraction between two polar molecules.
	+ In a dipole-dipole attraction, the partial positive side of one molecule is attracted to the partial negative side of an adjacent molecule. This is the strongest type of intermolecular force.
	+ Hydrogen bonds are among the strongest types of dipole-dipole forces. A hydrogen bond forms between a hydrogen atom in one molecule and a highly electronegative atom (such as oxygen) in another molecule, as shown at right.
* Dipole-induced dipole force – an intermolecular force of attraction that occurs when a polar molecule induces polarity in a nonpolar molecule.
	+ When the positive side of a polar molecule approaches a nonpolar molecule, the electron cloud of the nonpolar molecule becomes distorted, moving towards the polar molecule. As a result, the nonpolar molecule becomes temporarily polar, and is attracted to the polar molecule.
	+ A similar phenomenon occurs when the negative side of a polar molecule approaches a nonpolar molecule. Electrons are repelled from one side of the nonpolar molecule, causing it to become temporarily polar.
* Electronegativity – a measure of how strongly one atom attracts a shared pair of electrons. Electronegativity values range from 0.7 to 4.0, and are unitless.
	+ Nonmetals tend to have a greater electronegativity than metals.
	+ Electronegativity tends to increase across a period and decrease down a group in the periodic table.
* Intermolecular force – a force of attraction between molecules.
	+ Intermolecular forces (IMFs) are also known as van der Waals forces.
	+ Intermolecular forces are relatively weak compared to ionic and covalent bonds.
* Ionic bond – a bond formed by the attraction between two oppositely charged ions.
	+ Ionic bonds are the most polar type of bond since valence electrons are completely transferred from one atom to another.
	+ Ionic bonds occur if the electronegativity difference is greater than/equal to 1.67.
* London dispersion force – an intermolecular force of attraction that occurs when a temporary dipole of one atom induces polarity in a neighboring atom.
	+ Temporary dipoles form due to random fluctuations in the density of the electron cloud surrounding the nucleus.
	+ While London dispersion forces are the primary type of intermolecular forces between nonpolar molecules, they are also present between polar molecules.
	+ London dispersion forces are the weakest type of intermolecular force.
* Molecule – a stable particle composed of two or more atoms.
	+ A water molecule (H2O) is made of two hydrogen atoms and one oxygen atom.
* Nonpolar – having an overall even distribution of positive and negative charges.
	+ Both bonds and molecules as a whole can be nonpolar.
	+ Intermolecular forces between nonpolar molecules tend to be weaker than those between polar molecules.
* Nonpolar covalent bond – a covalent bond characterized by equal sharing of electrons.
	+ If the atoms in a bond have the same electronegativity (or a difference less than or equal to 0.4), then the bond is considered to be nonpolar covalent.
	+ All diatomic molecules (O2, Cl2, etc.) have nonpolar covalent bonds.
* Partial charge – a charge with an absolute value of less than 1 in regions of polar molecules created by unequal sharing of electron pairs in covalent bonds.
	+ Partial charges are represented by the Greek lowercase delta (δ+ or δ–).
	+ In a polar molecule, the region where the electron cloud is denser has a δ–charge while the region where the electron cloud is less dense has a δ+ charge.
* Polar – having distinct regions of positive and negative charge.
	+ Both individual bonds and molecules as a whole can be classified as polar.
	+ IMFs between polar molecules are relatively strong.
	+ A polar substance, such as water, will have both a partial positive and a partial negative side but overall will be neutral.

* Polar covalent bond – a covalent bond characterized by unequal sharing of electrons.
	+ If the atoms in a bond have an electronegativity difference greater than 0.4 and less than 1.67 then the bond is considered polar covalent.
	+ Each of the atoms in a polar covalent bond will either have a partial positive (δ+) or a partial negative (δ–) charge.
* Valence electron – an electron in the outermost energy level of an atom.
	+ Valence electrons take part in chemical bonds when they are transferred or shared.