**Vocabulary:** **Ionic Bonds**

dictionary2

**Vocabulary**

* Chemical family – a vertical column in the periodic table of elements.
  + Elements in the same family have the same number of valence electrons and have similar chemical properties.
* Ion – an electrically charged atom that has gained or lost one or more electrons.
  + An atom with more protons than electrons is a positively charged ion, or *cation*.
  + An atom with more electrons than protons is a negatively charged ion, or *anion*.
  + In an element symbol, the electric charge is shown at the upper right.
* Ionic bond – a bond formed by the attraction between two oppositely charged ions.
  + Positively charged ions attract negatively charged ions and vice versa.
  + Positively charged ions repel positively charged ions.
  + Negatively charged ions repel negatively charged ions.
* Ionization energy – the energy required to remove a valence electron from an atom.
  + The greater the ionization energy, the harder it is to remove an electron.
  + Metals are characterized by low ionization energy and tend to lose electrons easily. Nonmetals have high ionization energy and tend to gain electrons.
* Metal – an element that is malleable and usually conducts heat and electricity well.
  + Metal atoms tend to lose electrons when forming chemical bonds.
* Nonmetal – an element that is generally a poor conductor of heat and electricity.
  + Nonmetal atoms tend to gain or share electrons when forming chemical bonds.
  + Most nonmetals are gases at room temperature.
* Octet rule – a rule of thumb that states that atoms are most stable when surrounded by eight valence electrons.
  + Metals lose valence electrons to obtain a stable configuration.
  + Nonmetals gain or share electrons to obtain a stable configuration.
  + Elements with five or fewer electrons are exceptions to the octet rule because they become stable when they have two valence electrons.
* Shell – a particular region where electrons can orbit the nucleus of an atom.
* Valence electron – an electron in the outermost shell of an atom.