**Vocabulary:** **Solubility and Temperature**

dictionary2

**Vocabulary**

* Concentration – a measure of how much of a given substance is mixed with another substance.
  + To measure the concentration of a *solution*, divide the mass of the *solute* by the volume of the *solvent*.
  + A common unit of concentration is grams per 100 milliliters (g/100 mL) or grams per deciliter (g/dL).
* Dissolve – to pass into solution.
  + For example, sugar or salt can dissolve into water. The resulting solution is a *homogeneous mixture*.
* Homogeneous mixture – a combination of two or more substances that is exactly the same throughout.
  + Any small sample of a homogeneous mixture would look exactly the same as any other sample, even at a microscopic level.
  + Solutions are homogeneous mixtures.
* Solubility – the maximum concentration of solute that can be dissolved in a solvent.
  + Solubility depends on the type of solute and solvent used and often depends on the temperature and pressure of the solvent.
* Solubility curve – a graph showing the relationship between solubility and another variable such as temperature or pressure.
* Solute – a substance that is dissolved in another substance to form a solution.
  + In salt water, the solute is salt.
* Solution – a homogeneous mixture of two or more substances.
  + Solutions generally consist of a solute that is dissolved into a solvent.
    - Solvents are generally liquids.
    - Solutes can be solids, liquids, or gases.
  + Examples of solutions include salt water, sugar water, and seltzer.
* Solvent – a solid, liquid, or gas in which a solute is dissolved to form a solution.
  + In salt water, the solvent is water.